

Chapter 9 Chemical Names Formulas Answer Key

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Chapter 9: Chemical Names and Formulas. STUDY. PLAY. Monatomic Ion.
Consists of a single atom with a positive or negative charge resulting
from the loss of gain of one or more valence electrons and behaves as
a unit, yet it still has a charge. (+ or -). Determined by the amount
of electrons lost or gained.

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To name a binary molecular compound, use the following guidelines: 1.
Write the names of the elements in the order listed in the formula 2.
Use prefixes appropriately to indicate the number of each kind of atom
3. End the name of the second element with the suffix -ide

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First, confirm that the compound is a binary molecular compound. Name
the elements in the order listed in the formula. Use prefixes to
indicate the number of each kind of atom. Omit the prefix -mono when
the formula contains only one atom of the first element in the name.
The suffix of the name of the second element is -ide.

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1. when the name of the anion (X) ends in -ide the acid name begins
with the prefix hydro-. The stem of the anion has the suffix -ic and
is followed by the word acid (ex. HCl= hydrochloric acid) 2. when the
anion name ends in -ite the acid name is the stem of the anion with
the suffix -ous followed by the word acid (ex. H₂SO₃ -- X=sulfite =
sulfurous acid)

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Writing the Names and Formulas for Bases. A base is a compound that
produces OH⁻ in water. When naming a base, you name it like any other
compound that starts with a regular or transition metal. Ex: NaOH =
sodium hydroxide. When writing the formula for a base, remember to
balance charges. Ex: magnesium hydroxide = Mg(OH) ₂

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Chapter 9 Chemical Names and Formulas 83 SECTION 9.3 NAMING AND WRITING
FORMULAS FOR MOLECULAR COMPOUNDS (pages 268-270) This section explains
the rules for naming and writing formulas for binary molecular
compounds. Naming Binary Molecular Compounds (pages 268-269) 1. Circle
the letter of the type(s) of elements that form binary molecular

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compounds.

~~Name Date Class CHEMICAL NAMES AND FORMULAS 9~~

SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC COMPOUNDS I. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3. Name the following binary ionic compounds. sodium sulfide g. CuCl : h. SnCl_2 . a. MnO_2 c. d. SrBr_2 221 Chapter 9 Chemical Names and Formulas

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Chemical Names and Formulas 281 CHAPTER 9 Assessment 42. a. 2^- b. 1^+ c. 1^- d. 3^+ 43. a. 2^+ b. 2^+ c. 3^+ d. 1^+ 44. a. barium ion b. iodide ion c. silver ion d. mercury(II) ion 45. cyanide, CN^- and hydroxide, OH^- 46. a. hydroxide ion b. lead(IV) ion c. sulfate ion d. oxide ion 47. zero; A compound is electrically neutral. 48. The symbols for the cation and

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Chapter 9 Chemical Names Formulas Chapter 9 - Chemical Names and Formulas. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. bibiloni1901. Terms in this set (20) Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence

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Chapter 9 "Chemical Names and Formulas" Tools. Copy this to my account; E-mail to a friend; Find other activities; Start over; Help; Use these activities to help yourself to learn the vocabulary and major topics presented in this chapter. A B; anion: any atom or group of atoms that has a negative charge:

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Chapter 9 Practice Test - Naming and Writing Chemical Formulas Answer Section MATCHING 1. ANS: A PTS: 1 DIF: L1 REF: p. 253 OBJ: 9.1.1 Identify the charges of monatomic ions by using the periodic table, and name the ions. STA: Ch.5.a 2.ANS: H PTS: 1 DIF: L1 REF: p. 254

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